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**[1-(4-Hydroxyphenyl)-1*H*-tetrazol-5-yl-sulfanyl]acetic acid**

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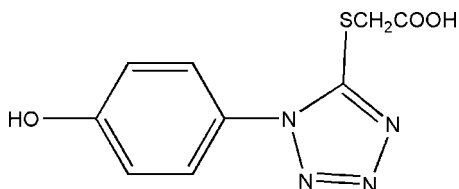
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Key indicators: single-crystal X-ray study;  $T = 293$  K; mean  $\sigma(\text{C}-\text{C}) = 0.003$  Å;  $R$  factor = 0.047;  $wR$  factor = 0.106; data-to-parameter ratio = 15.9.

The title compound,  $\text{C}_9\text{H}_8\text{N}_4\text{O}_3\text{S}$ , shows a layer structure constructed from intermolecular  $\text{O}-\text{H}\cdots\text{O}$  and  $\text{O}-\text{H}\cdots\text{N}$  hydrogen bonds. Interatomic distances suggest that extensive, but not uniform,  $\pi$ -electron delocalization is present in the tetrazole rings and extends over the exocyclic  $\text{C}-\text{S}$  bond.

## Related literature

For related literature on tetrazol-5-thione and its derivatives, see: Cea-Olivares *et al.* (1997); Kim *et al.* (2003).



## Experimental

## Crystal data

$\text{C}_9\text{H}_8\text{N}_4\text{O}_3\text{S}$   
 $M_r = 252.25$

Orthorhombic,  $Pbca$   
 $a = 14.407$  (3) Å

$b = 7.3365$  (16) Å  
 $c = 21.107$  (5) Å  
 $V = 2231.0$  (9) Å<sup>3</sup>  
 $Z = 8$

Mo  $K\alpha$  radiation $\mu = 0.29$  mm<sup>-1</sup> $T = 293$  K $0.28 \times 0.16 \times 0.10$  mm

## Data collection

Bruker APEXII area-detector diffractometer  
 Absorption correction: multi-scan (*SADABS*; Sheldrick, 1996)  
 $T_{\min} = 0.95$ ,  $T_{\max} = 0.97$

18595 measured reflections  
 2550 independent reflections  
 1834 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.099$

## Refinement

$R[F^2 > 2\sigma(F^2)] = 0.047$   
 $wR(F^2) = 0.106$   
 $S = 1.06$   
 2550 reflections  
 160 parameters  
 2 restraints

H atoms treated by a mixture of independent and constrained refinement

 $\Delta\rho_{\text{max}} = 0.19$  e Å<sup>-3</sup> $\Delta\rho_{\text{min}} = -0.21$  e Å<sup>-3</sup>

Table 1

Hydrogen-bond geometry (Å, °).

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
$\text{O1}-\text{H1}\cdots\text{O3}^{\text{i}}$	0.835 (17)	1.953 (17)	2.787 (3)	177 (3)
$\text{O2}-\text{H2}\cdots\text{N4}^{\text{ii}}$	0.834 (17)	1.866 (17)	2.699 (3)	176 (3)
$\text{O2}-\text{H2}\cdots\text{N3}^{\text{ii}}$	0.834 (17)	2.60 (2)	3.369 (3)	154 (3)

Symmetry codes: (i)  $x - \frac{1}{2}, y, -z + \frac{1}{2}$ ; (ii)  $-x + \frac{3}{2}, y - \frac{1}{2}, z$ .

Data collection: *APEX2* (Bruker, 2002); cell refinement: *SAINT* (Bruker, 2002); data reduction: *SAINT*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *SHELXTL* (Sheldrick, 2008); software used to prepare material for publication: *SHELXL97*.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: AT2673).

## References

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**supplementary materials**

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## [1-(4-Hydroxyphenyl)-1*H*-tetrazol-5-ylsulfanyl]acetic acid

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### Comment

Tetrazol-5-thione and its derivatives are interesting ligands from a structural point of view since they can display a wide range of coordination patterns with metal ions. Due to a variety of potential coordination sites, they can act as monodentate (–S or –N) or bidentate (–N, N or –N, S) ligands, forming polymers or interacting with metal ions (Cea-Olivares *et al.*, 1997; Kim *et al.*, 2003).

As shown in Fig.1, the bond lengths within the tetrazole ring exhibit the expected pattern of four long bonds (C7—N1, C7—N4, N3—N4 and N1—N2) together with a short one (N2—N3). In detail, C7—N1 [1.340 (2) Å] and C7—N4 [1.324 (2) Å] are typical for carbon-nitrogen single bonds from 1.336 Å to 1.420 Å, while N3—N4 [1.365 (2) Å] and N1—N2 [1.358 (2) Å] are between the single and double bonds. And the bond distance N2—N3 of 1.283 (2) Å is similar to that of a double bond of 1.25 Å. The bond length of S1—C7 [1.723 (2) Å] also falls between the double and single bonds. All these interatomic distances suggest that extensive but not uniform  $\pi$  electron delocalization is present in the tetrazole rings and extends over the exocyclic C—S bond.

### Experimental

To an aqueous solution of 1-(4-hydroxyphenyl)-5-thiotetrazole (1.940 g, 10.0 mmol) and NaOH (0.80 g, 20.0 mmol) were sequentially added the aqueous solution of chloroacetic acid (2.835 g, 30.0 mmol) and NaOH (1.400 g, 35.0 mmol). After stirring for 4 h at 353 K under nitrogen atmosphere, the mixture was cooled to room temperature slowly. Adjusted the pH to 2 by adding 1.0 mol/L HCl, the white deposit appeared rapidly. The solids were filtered and washed with water. The single crystals suitable for X-ray diffraction were obtained by the re-crystallization of sieved solid in the ethanol.

### Refinement

The H atoms bonded to C atoms were positioned geometrically and treated as riding, [aromatic C—H = 0.93 Å and aliphatic C—H = 0.97 Å,  $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C})$ ]. The H atoms bonded to O atoms were located in a difference Fourier map and refined isotropically.

### Figures

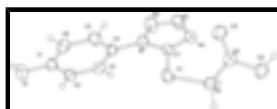


Fig. 1. The molecular structure of the title compound, showing 30% probability displacement ellipsoids

## [1-(4-Hydroxyphenyl)-1H-tetrazol-5-ylsulfanyl]acetic acid

### Crystal data

$C_9H_8N_4O_3S$	$F_{000} = 1040$
$M_r = 252.25$	$D_x = 1.502 \text{ Mg m}^{-3}$
Orthorhombic, <i>Pbca</i>	Mo $K\alpha$ radiation
Hall symbol: -P 2ac 2ab	$\lambda = 0.71073 \text{ \AA}$
$a = 14.407 (3) \text{ \AA}$	Cell parameters from 2531 reflections
$b = 7.3365 (16) \text{ \AA}$	$\theta = 2.4\text{--}27.5^\circ$
$c = 21.107 (5) \text{ \AA}$	$\mu = 0.29 \text{ mm}^{-1}$
$V = 2231.0 (9) \text{ \AA}^3$	$T = 293 \text{ K}$
$Z = 8$	Block, colourless
	$0.28 \times 0.16 \times 0.10 \text{ mm}$

### Data collection

Bruker APEXII area-detector diffractometer	2550 independent reflections
Radiation source: fine-focus sealed tube	1834 reflections with $I > 2\sigma(I)$
Monochromator: graphite	$R_{\text{int}} = 0.099$
$T = 293 \text{ K}$	$\theta_{\text{max}} = 27.5^\circ$
$\omega$ scans	$\theta_{\text{min}} = 2.4^\circ$
Absorption correction: multi-scan (SADABS; Sheldrick, 1996)	$h = -18 \rightarrow 18$
$T_{\text{min}} = 0.95, T_{\text{max}} = 0.97$	$k = -9 \rightarrow 9$
18595 measured reflections	$l = -27 \rightarrow 27$

### Refinement

Refinement on $F^2$	Secondary atom site location: difference Fourier map
Least-squares matrix: full	Hydrogen site location: inferred from neighbouring sites
$R[F^2 > 2\sigma(F^2)] = 0.047$	H atoms treated by a mixture of independent and constrained refinement
$wR(F^2) = 0.106$	$w = 1/[\sigma^2(F_o^2) + (0.0144P)^2 + 1.7635P]$
$S = 1.06$	where $P = (F_o^2 + 2F_c^2)/3$
2550 reflections	$(\Delta/\sigma)_{\text{max}} = 0.001$
160 parameters	$\Delta\rho_{\text{max}} = 0.19 \text{ e \AA}^{-3}$
2 restraints	$\Delta\rho_{\text{min}} = -0.21 \text{ e \AA}^{-3}$
Primary atom site location: structure-invariant direct methods	Extinction correction: none

*Special details*

**Geometry.** All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > \sigma(F^2)$  is used only for calculating  $R$ -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ -factors based on ALL data will be even larger.

*Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )*

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
S1	0.51454 (4)	0.15070 (9)	0.09961 (3)	0.05074 (19)
O1	0.15861 (13)	0.3256 (3)	0.28120 (9)	0.0610 (5)
H1	0.177 (2)	0.270 (4)	0.3133 (11)	0.073*
O2	0.75012 (12)	0.0102 (3)	0.01914 (8)	0.0551 (5)
H2	0.8048 (13)	0.010 (4)	0.0322 (13)	0.066*
O3	0.72494 (12)	0.1471 (3)	0.11139 (8)	0.0645 (5)
N1	0.45507 (13)	0.4915 (3)	0.12461 (10)	0.0468 (5)
N2	0.47350 (16)	0.6679 (3)	0.10961 (12)	0.0637 (6)
N3	0.54403 (16)	0.6653 (3)	0.07237 (12)	0.0619 (6)
N4	0.57366 (13)	0.4912 (3)	0.06208 (10)	0.0480 (5)
C1	0.23393 (16)	0.3617 (3)	0.24421 (11)	0.0441 (5)
C2	0.21686 (16)	0.4284 (3)	0.18436 (12)	0.0470 (6)
H2A	0.1561	0.4457	0.1708	0.056*
C3	0.28941 (16)	0.4695 (3)	0.14462 (12)	0.0478 (6)
H3A	0.2783	0.5152	0.1042	0.057*
C4	0.37920 (15)	0.4418 (3)	0.16560 (11)	0.0436 (5)
C5	0.39734 (17)	0.3749 (4)	0.22511 (12)	0.0509 (6)
H5A	0.4582	0.3576	0.2386	0.061*
C6	0.32408 (17)	0.3336 (4)	0.26463 (12)	0.0509 (6)
H6A	0.3352	0.2870	0.3049	0.061*
C7	0.51693 (14)	0.3855 (3)	0.09473 (11)	0.0402 (5)
C8	0.59818 (15)	0.0925 (3)	0.03998 (12)	0.0465 (6)
H8A	0.5826	-0.0265	0.0231	0.056*
H8B	0.5931	0.1797	0.0056	0.056*
C9	0.69718 (16)	0.0887 (3)	0.06206 (11)	0.0429 (5)

*Atomic displacement parameters ( $\text{\AA}^2$ )*

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
S1	0.0418 (3)	0.0418 (3)	0.0686 (4)	-0.0030 (3)	0.0098 (3)	-0.0001 (3)
O1	0.0472 (10)	0.0753 (14)	0.0607 (11)	0.0032 (9)	0.0128 (9)	0.0151 (10)
O2	0.0389 (8)	0.0757 (13)	0.0508 (10)	0.0116 (9)	-0.0034 (8)	-0.0121 (9)

## supplementary materials

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O3	0.0490 (11)	0.0903 (15)	0.0542 (11)	0.0078 (10)	-0.0099 (8)	-0.0217 (11)
N1	0.0421 (11)	0.0421 (12)	0.0563 (12)	0.0009 (9)	0.0080 (9)	0.0042 (10)
N2	0.0611 (14)	0.0407 (12)	0.0894 (17)	0.0025 (11)	0.0171 (13)	0.0099 (12)
N3	0.0520 (13)	0.0486 (13)	0.0850 (16)	-0.0031 (11)	0.0140 (12)	0.0113 (12)
N4	0.0375 (10)	0.0481 (12)	0.0584 (12)	-0.0027 (9)	0.0030 (9)	0.0048 (10)
C1	0.0426 (13)	0.0412 (13)	0.0484 (13)	0.0017 (10)	0.0062 (10)	-0.0012 (11)
C2	0.0389 (12)	0.0491 (14)	0.0529 (14)	0.0016 (11)	-0.0015 (11)	0.0025 (12)
C3	0.0486 (14)	0.0487 (15)	0.0460 (13)	0.0035 (11)	-0.0012 (11)	0.0050 (11)
C4	0.0414 (12)	0.0392 (12)	0.0502 (14)	0.0007 (10)	0.0080 (11)	0.0005 (11)
C5	0.0391 (12)	0.0588 (16)	0.0547 (15)	0.0029 (11)	-0.0014 (11)	0.0016 (12)
C6	0.0515 (14)	0.0555 (15)	0.0456 (14)	0.0018 (12)	0.0002 (11)	0.0041 (12)
C7	0.0319 (11)	0.0428 (13)	0.0458 (12)	-0.0020 (9)	0.0000 (10)	0.0000 (10)
C8	0.0394 (12)	0.0443 (14)	0.0558 (14)	0.0005 (10)	-0.0043 (11)	-0.0080 (11)
C9	0.0408 (12)	0.0445 (13)	0.0435 (13)	0.0040 (10)	-0.0030 (11)	0.0003 (11)

### *Geometric parameters (Å, °)*

S1—C7	1.726 (2)	C1—C2	1.377 (3)
S1—C8	1.794 (2)	C1—C6	1.384 (3)
O1—C1	1.363 (3)	C2—C3	1.374 (3)
O1—H1	0.835 (17)	C2—H2A	0.9300
O2—C9	1.317 (3)	C3—C4	1.382 (3)
O2—H2	0.834 (17)	C3—H3A	0.9300
O3—C9	1.195 (3)	C4—C5	1.374 (3)
N1—C7	1.341 (3)	C5—C6	1.379 (3)
N1—N2	1.359 (3)	C5—H5A	0.9300
N1—C4	1.441 (3)	C6—H6A	0.9300
N2—N3	1.285 (3)	C8—C9	1.501 (3)
N3—N4	1.364 (3)	C8—H8A	0.9700
N4—C7	1.321 (3)	C8—H8B	0.9700
C7—S1—C8	100.49 (11)	C3—C4—N1	118.7 (2)
C1—O1—H1	108 (2)	C4—C5—C6	119.1 (2)
C9—O2—H2	109 (2)	C4—C5—H5A	120.5
C7—N1—N2	108.24 (19)	C6—C5—H5A	120.5
C7—N1—C4	129.8 (2)	C5—C6—C1	119.8 (2)
N2—N1—C4	121.94 (19)	C5—C6—H6A	120.1
N3—N2—N1	106.4 (2)	C1—C6—H6A	120.1
N2—N3—N4	111.0 (2)	N4—C7—N1	108.4 (2)
C7—N4—N3	105.86 (19)	N4—C7—S1	129.03 (18)
O1—C1—C2	116.9 (2)	N1—C7—S1	122.53 (17)
O1—C1—C6	122.7 (2)	C9—C8—S1	115.16 (17)
C2—C1—C6	120.4 (2)	C9—C8—H8A	108.5
C3—C2—C1	120.2 (2)	S1—C8—H8A	108.5
C3—C2—H2A	119.9	C9—C8—H8B	108.5
C1—C2—H2A	119.9	S1—C8—H8B	108.5
C2—C3—C4	119.0 (2)	H8A—C8—H8B	107.5
C2—C3—H3A	120.5	O3—C9—O2	124.2 (2)
C4—C3—H3A	120.5	O3—C9—C8	125.6 (2)
C5—C4—C3	121.6 (2)	O2—C9—C8	110.2 (2)

C5—C4—N1	119.7 (2)		
C7—N1—N2—N3	-0.2 (3)	C4—C5—C6—C1	-0.7 (4)
C4—N1—N2—N3	-179.5 (2)	O1—C1—C6—C5	-179.3 (2)
N1—N2—N3—N4	-0.2 (3)	C2—C1—C6—C5	0.9 (4)
N2—N3—N4—C7	0.5 (3)	N3—N4—C7—N1	-0.6 (3)
O1—C1—C2—C3	179.4 (2)	N3—N4—C7—S1	177.64 (19)
C6—C1—C2—C3	-0.7 (4)	N2—N1—C7—N4	0.5 (3)
C1—C2—C3—C4	0.4 (4)	C4—N1—C7—N4	179.7 (2)
C2—C3—C4—C5	-0.2 (4)	N2—N1—C7—S1	-177.85 (18)
C2—C3—C4—N1	-177.7 (2)	C4—N1—C7—S1	1.3 (3)
C7—N1—C4—C5	70.7 (3)	C8—S1—C7—N4	-7.9 (2)
N2—N1—C4—C5	-110.2 (3)	C8—S1—C7—N1	170.13 (19)
C7—N1—C4—C3	-111.7 (3)	C7—S1—C8—C9	86.2 (2)
N2—N1—C4—C3	67.4 (3)	S1—C8—C9—O3	-12.3 (4)
C3—C4—C5—C6	0.3 (4)	S1—C8—C9—O2	167.39 (17)
N1—C4—C5—C6	177.9 (2)		

Hydrogen-bond geometry (Å, °)

<i>D</i> —H... <i>A</i>	<i>D</i> —H	H... <i>A</i>	<i>D</i> ... <i>A</i>	<i>D</i> —H... <i>A</i>
O1—H1...O3 <sup>i</sup>	0.835 (17)	1.953 (17)	2.787 (3)	177 (3)
O2—H2...N4 <sup>ii</sup>	0.834 (17)	1.866 (17)	2.699 (3)	176 (3)
O2—H2...N3 <sup>ii</sup>	0.834 (17)	2.60 (2)	3.369 (3)	154 (3)

Symmetry codes: (i)  $x-1/2, y, -z+1/2$ ; (ii)  $-x+3/2, y-1/2, z$ .

Fig. 1

